

Chapter 10 Study Guide - KEY

1. What is the molarity of a 2.05 L solution that is made by dissolving 27.5 g of sodium chloride in water?

$$\frac{27.5 \text{ g NaCl}}{58.5 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.470 \text{ mol}$$

$$M = \frac{\text{mol solute}}{\text{L sol'n}} = \frac{0.470 \text{ mol}}{2.05 \text{ L}} = \underline{0.229 M}$$

2. What is the normality of a 6.0 M solution of phosphoric acid?

$$N = (M)(\text{total pos. ox}) = (6.0 M)(3) = \underline{18 N}$$

3. What is the molality of a solution that is made by dissolving 98.7 g of calcium chloride in 755 g of water?

$$\frac{98.7 \text{ g CaCl}_2}{111.1 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.888 \text{ mol}$$

$$m = \frac{\text{mol solute}}{\text{kg solvent}} = \frac{0.888 \text{ mol}}{0.755 \text{ kg}} = \underline{1.18 m}$$

4. How many grams of KI are needed to prepare 325 mL of a 1.33 M solution?

$$1.33 M = \frac{n \text{ mol}}{0.325 \text{ L}}$$

$$n = 0.432 \text{ mol KI}$$

$$\text{mass} = \frac{0.432 \text{ mol KI}}{1 \text{ mol}} \times \frac{166.1 \text{ g}}{1} = \underline{71.8 \text{ g KI}}$$

5. How many grams of NaCl are needed to prepare a 4.15 m solution, using 176 g of water?

$$4.15 m = \frac{n \text{ mol}}{0.176 \text{ kg}}$$

$$n = 0.7304 \text{ mol NaCl}$$

$$\text{mass} = \frac{0.7304 \text{ mol NaCl}}{1 \text{ mol}} \times \frac{58.5 \text{ g}}{1} = \underline{42.7 \text{ g NaCl}}$$

6. Ethyl iodide ($\text{C}_2\text{H}_5\text{I}$) boils at 72.5°C and has a density of 1.933 g/mL .

- a. A solution prepared by dissolving 0.300 mol of a nonelectrolyte in 750.0 mL of ethyl iodide boils at 73.5°C . What is the boiling point constant (K_b) for ethyl iodide?

$$\text{C}_2\text{H}_5\text{I} = 156.0 \text{ g/mol}$$

$$\text{mass solvent} = \frac{750.0 \text{ mL}}{1 \text{ mL}} \times \frac{1.933 \text{ g}}{1} = 1449.75 \text{ g solvent}$$

$$m = \frac{0.300 \text{ mol}}{1.44975 \text{ kg}} = 0.207 m$$

$$\Delta T_b = K_b \cdot m$$

$$1.0^\circ\text{C} = (K_b)(0.207 m)$$

$$K_b = \underline{4.83^\circ\text{C/m}}$$

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