

Electronegativity and Bond Polarity

Electronegativity

We have discussed chemical bonding in terms of two models: ionic and covalent bonding. In doing so, we have assumed that the bond between two atoms is always exclusively one or the other. Looking at the Lewis Structure for hydrochloric acid, this implies that the pair of electrons shared by the covalent bond is shared equally. However, this is not the case.



On theoretical grounds, chemists believe that atoms have different abilities to attract electrons. For example, the farther away from the nucleus the electrons are, the weaker their attraction is to the nucleus. Also, inner electrons shield the valence electrons from the attraction of the positive nucleus.

multiple
electron
shells

Chemists use the term electronegativity to describe the relative ability of an atom to attract a pair of bonding electrons in its valence level. The higher the electronegativity of an element, the more that atom will attempt to pull electrons towards itself (and away from any atom it bonds to). Electronegativity is usually assigned on a scale developed by Linus Pauling. Pauling based his scale on energy changes in chemical reactions.

To determine the electronegativity value for an element, refer to your periodic table. There is a general trend that you can see from your periodic table: as you move from left to right and bottom to top, the electronegativity increases. According to the scale, Fluorine has the highest electronegativity, 4.0, and Francium has the lowest electronegativity, 0.7. As a general rule, metals tend to have low electronegativities while non-metals tend to have high electronegativities.



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Susan McMurry, John McMurry



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