

STUDY GUIDE FOR MODULE NO. 8

CHEMICAL KINETICS



MODULE OVERVIEW

Chemical reactions vary not only with the products that they produce, but also how fast or slow they occur. Some chemical reactions occur slowly, such as the formation of rust while others, like the explosion of fireworks, occur quickly and instantaneously. For this Module, we would be concerned with chemical kinetics, the area of chemistry which concerns reaction rates.



MODULE LEARNING OBJECTIVES

At the end of this Module, you should be able to:

1. explain how certain factors affect rates of reactions
2. determine the rate law of reactions
3. use the Arrhenius equation to relate rate constant to temperature and activation energy
4. identify the factors affecting effective collisions
5. Differentiate homogeneous and heterogeneous catalysis
6. sketch a potential energy profile showing activation energies for the forward and reverse reactions, and how they are affected by the addition of a catalyst



FACTORS AFFECTING REACTION RATES

Four factors affect the rate at which any particular reaction occurs:

1. **Physical state of reactants.** For a reaction to occur, the reactants should come in contact. The more opportunities for contact, the faster the reactions may occur. For heterogeneous reactions (those reactions where the reactants are of different phases), the rate is increased when the reactants have more areas of contact.
2. **Concentration of reactants.** Generally, the higher the concentration of reactants, the faster that a reaction occurs. This is because the more concentrated the reactants are, the more opportunities that they have to collide, resulting to a faster rate.
3. **Temperature.** You have learned from previous Modules that the higher the temperature, the higher the kinetic energy of the particles, resulting to more rapid movements and increased collision. Because of this, reactions occur more rapidly.
4. **Presence of a catalyst.** A catalyst is a substance which increases the rate of a certain chemical reaction but is not used up. When a catalyst is used to speed up a reaction, the collisions that occur between the reactants are affected, which leads to a faster rate.

Why these factors affect reaction rate will be discussed in more detail as we move on with this module.



LEARNING ACTIVITY 1

Cite at least three (3) examples or instances where the each of the given factors can be observed affecting reaction rates. Justify your answers.



REACTION RATES

Before delving into the next parts of the module, you may want to watch the video clip *Introduction to Kinetics / Energy and enzymes / Biology / Khan Academy* using the link https://www.youtube.com/watch?v=___zy-oOLPug&list=PL166048DD75B05C0D&index=40. This will help you to have an background on the concept of kinetics and the connections of each topic for this module. You may also want to refer to this video as you move on with the module to have a more visual and deep understanding of reaction rates.



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Table of Contents Study Guide Chemical Kinetics

1. Understanding the eBook Study Guide Chemical Kinetics
 - The Rise of Digital Reading Study Guide Chemical Kinetics
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Chemical Kinetics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Chemical Kinetics
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Chemical Kinetics
 - Personalized Recommendations
 - Study Guide Chemical Kinetics User Reviews and Ratings
 - Study Guide Chemical Kinetics and Bestseller Lists
5. Accessing Study Guide Chemical Kinetics Free and Paid eBooks
 - Study Guide Chemical Kinetics Public Domain eBooks
 - Study Guide Chemical Kinetics eBook Subscription Services
 - Study Guide Chemical Kinetics Budget-Friendly Options
6. Navigating Study Guide Chemical Kinetics eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Chemical Kinetics Compatibility with Devices
 - Study Guide Chemical Kinetics Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Chemical Kinetics
 - Highlighting and Note-Taking Study Guide Chemical Kinetics
 - Interactive Elements Study Guide Chemical Kinetics
8. Staying Engaged with Study Guide Chemical Kinetics

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Chemical Kinetics
- 9. Balancing eBooks and Physical Books Study Guide Chemical Kinetics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Chemical Kinetics
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Chemical Kinetics
 - Setting Reading Goals Study Guide Chemical Kinetics
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Chemical Kinetics
 - Fact-Checking eBook Content of Study Guide Chemical Kinetics
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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