

Introduction

The Oxford University Style Guide aims to provide a guide to writing and formatting documents written by staff on behalf of the University (or one of its constituent departments etc). It is part of the University's branding toolkit (www.ox.ac.uk/branding_toolkit) which enables the University's formal documentation to be presented consistently across all communications.

The style guide is not intended for public or external use, and does not purport to compete with OUP's professional writing guides and dictionaries.

Objectives of the style guide

We have three main objectives in writing this style guide:

- to provide an all-purpose guide to consistent presentation for University staff in written communications;
- to review the guide at least once a year, ensuring that it properly reflects modern usage and is fit for purpose, and to update it as required;
- as part of the review process, to invite proposals from members of the University who disagree with any existing guidance, and to act as an arbiter on those cases.

How the guide is arranged

The style guide is intended to be read as an interactive PDF, where it can be cross-referenced. However, the PDF can be printed if preferred for ease of reference.

If we update the style guide we will highlight on the main webpage (www.ox.ac.uk/styleguide) whether anything has changed as well as changing the term listed on the front cover.

How to use the guide

- search for a specific term (such as **semicolon**)
- browse through a section (such as **Punctuation**)

What is/is not included in the style guide

The guide does not tell you how to write. We aim to help you write correctly, and to encourage consistency across the University's written communications.

Quick reference guide

The general rule

If there are multiple (correct) ways of doing something, choose the one which uses the least space and the least ink. For instance:

- close up spaces and don't use full stops in abbreviations (eg *figs*)
- use lower case wherever possible
- only write out numbers up to ten and use figures for 11 onwards.

University of Oxford or Oxford University?

These terms are interchangeable and can either be alternated for variety or kept the same for consistency.

University branding information

Other information on University branding, including the use of the logo, can be found online at www.ox.ac.uk/branding_toolkit

Queries

If you have any queries about using this guide, please contact:

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The Oxford Guide To Style

John R. Kohl



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the objectives of the boiling point determination lab were to determine the boiling point of a mystery substance provided to each student boiling point is the point at which a liquid changes physical states and becomes a vapor or gas boiling point is important to understand because it can be used to help distinguish between different substances

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise

calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affect by three key factors the first being intermolecular forces inter molecular affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in applied pressure so all boiling points should be reported with the measured pressure a compound's normal boiling point refers to its boiling point at a pressure of 760 mm Hg

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 °C so in order to convert an experimental measurement taken in Calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that needs to be determined by physical properties such as boiling point melting point density solubility etc in this experiment I aim to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a Bunsen burner boiling

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there are a variety of methods by which a sample's boiling point can be determined including distillation reflux and by using a thiele tube the most straightforward method uses a thiele tube and has the advantage of using less than 0.5 ml of material

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