STONY BROOK UNIVERSITY CHE 132: General Chemistry II Workshop K_{ap}

A. Solubility Product

- 1. What is the molar solubility of the following compounds:
 - K₁₀ = 1.77 x 10⁻¹⁰ a. AgCl (aq)

b. MgF₂ (aq) $K_{\rm ap} = 5.16 \times 10^{-11}$

c. lead(II) phosphate(aq)

$$K_{top} = 1.0 \times 10^{-54}$$

- Effect of common ions and pH on K_{sp}.
 Consider calcium carbonate CaCO₃ (limestone) with a K_{sp} of 4.96 x 10⁻⁹. What is the molar solubility of CaCO₃?

b. What is the molar solubility of CaCO₃ at pH 5.0?

c. What is the molar solubility of CaCO₃ in a solution of 0.100 M NaCO₃(aq)?

Stony Brook Che 134 Lab Manual

Daniel F McAuley

Stony Brook Che 134 Lab Manual:

Stony Brook Che 134 Lab Manual Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the power of words has are more evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **Stony Brook Che 134 Lab Manual**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall affect readers.

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