

Stoichiometry Review

Directions: Each of the problems below reviews a different type of stoichiometry. Be sure to read the problem carefully to know what it is asking for.

Mass-to-Mass Problems

1. How many grams of potassium chloride, KCl, are produced during the decomposition of 100.0 grams of potassium chlorate, KClO_3 ?



$$\frac{100 \text{ g KClO}_3}{1} \times \frac{1 \text{ mol KClO}_3}{122.545 \text{ g KClO}_3} \times \frac{1 \text{ mol KCl}}{1 \text{ mol KClO}_3} \times \frac{74.548 \text{ g KCl}}{1 \text{ mol KCl}} = \underline{60.83 \text{ g KCl}}$$

2. Find the number of grams of HCl needed to react completely with 0.50 moles of magnesium.



$$\frac{0.5 \text{ mol Mg}}{1} \times \frac{1 \text{ mol HCl}}{1 \text{ mol Mg}} \times \frac{36.458 \text{ g HCl}}{1 \text{ mol HCl}} = \underline{18.23 \text{ g HCl}}$$

3. Find the mass in grams of hydrogen gas produced when 14.0 moles of HCl are added to an excess amount of magnesium.



$$\frac{14 \text{ mol HCl}}{1} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} = \underline{14.112 \text{ g H}_2}$$

4. If 0.75 moles of ammonia are needed, how many grams of nitrogen will be consumed?



$$\frac{0.75 \text{ mol NH}_3}{1} \times \frac{1 \text{ mol N}_2}{2 \text{ mol NH}_3} \times \frac{28.014 \text{ g N}_2}{1 \text{ mol N}_2} = \underline{10.51 \text{ g N}_2}$$

Stoichiometry Review Packet

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Stoichiometry Review Packet:

Stoichiometry Unit Project Luann Marie Decker,1998 *The Big Book of Chemistry Teacher Stories* Jeff Lark, Stories from years of teaching high school chemistry **Resources in Education** ,1980 **Scientific and Technical Aerospace Reports** ,1994 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database *Basic Chemistry* William S. Seese,Guido H. Daub,1981 Chemistry John Olmsted,Gregory M. Williams,1997 Textbook outlining concepts of molecular science Science and Technology of Crystal Growth J.P. van der Eerden,O.S.L. Bruinsma,2012-12-06

1 The ninth International Summer School on Crystal Growth ISSCG IX A complete theory of crystal growth establishes the full dependence of crystal size shape and structure on external parameters like temperature pressure composition purity growth rate and stirring of the mother phase implicitly establishing how the corresponding fields vary in space and time Such a theory does not exist however Therefore equipment to grow crystals is developed on the basis of partial knowledge Skill experience and creativity still are of central importance for the success o a crystal growth system In this book we collected contributions from the teachers of the ninth International Summer School on Crystal Growth ISSCG IX held 11 16 june 1995 at Papendal the national sports centre of the Netherlands These contributions were used during the lectures The authors have tried to present their work in such a way that only basic physical knowledge is required to understand the papers The book can be used as an introduction to various important sub disciplines of the science and technology of crystal growth Since however the information content considerably exceeds a lecture note level and touches the present limits of understanding it is an up to date handbook as well **Bell Laboratories Talks and Papers** Bell Telephone Laboratories. Libraries and Information Systems Center,1982 Bibliography of Agriculture ,1990 **Phase Diagrams for Ceramists** Ernest M. Levin,Carl R. Robbins,Howard F. McMurdie,1964

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