

Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 9: Solving Equilibrium Problems for Complex Systems

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 9

- 9-2.** Why are simplifying assumptions in equilibrium problems restricted to relationships that are sums or differences?

Answer:

To simplify equilibrium calculations, we sometimes assume that the concentrations of one or more species are negligible and can be approximated as 0.00 M. In a sum or difference assuming a concentration is 0.00 M leads to an appropriate result. In contrast, if we were to simplify an equilibrium constant expression by assuming one or more concentrations are zero, we would be multiplying or dividing by 0.00, which would render the expression meaningless.

- 9-4.** Why do molar concentrations of some species appear as multiples in charge-balance equations?

Answer:

A charge-balance equation is derived by relating the concentration of cations and anions
no. mol/L positive charge = no. mol/L negative charge

For a doubly charged ion, such as Ba^{2+} , the concentration of charge for each mole is twice the molar concentration of the Ba^{2+} . That is,

$$\text{mol/L positive charge} = 2[\text{Ba}^{2+}]$$

Thus, the molar concentration of all multiply charged species is always multiplied by the charge in a charge-balance equation.

- 9-5.** Write the mass-balance expressions for a solution that is

Solutions:

- (a) 0.2 M in HF.

$$0.20 = [\text{HF}] + [\text{F}^-]$$

- (c) 0.10 M in H_3PO_4 .

$$0.10 = [\text{H}_3\text{PO}_4] + [\text{H}_2\text{PO}_4^-] + [\text{HPO}_4^{2-}] + [\text{PO}_4^{3-}]$$

- (e) 0.0500 M in HClO_2 and 0.100 M in NaClO_2 .

$$0.0500 + 0.100 = [\text{HClO}_2] + [\text{ClO}_2^-]$$

$$[\text{Na}^+] = C_{\text{NaClO}_2} = 0.100 \text{ M}$$

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Seppo Sorvari, Anna Maria Pirttilä



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26 Molecular Absorption Spectroscopy 211 9 27 Molecular Fluorescence Spectroscopy 220 3 28 Atomic Spectroscopy 223 5
29 Kinetic Methods of Analysis 228 6 30 An Introduction to Analytical Separations 234 7 31 Gas Chromatography 241 3 32
High Performance Liquid Chromatography 244 3 33 Miscellaneous Separation Methods 247 2 35 Preparing Samples for
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1. Understanding the eBook Solution Manual Skoog
 - The Rise of Digital Reading Solution Manual Skoog
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Skoog
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Skoog
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Skoog
 - Personalized Recommendations
 - Solution Manual Skoog User Reviews and Ratings
 - Solution Manual Skoog and Bestseller Lists
5. Accessing Solution Manual Skoog Free and Paid eBooks
 - Solution Manual Skoog Public Domain eBooks
 - Solution Manual Skoog eBook Subscription Services
 - Solution Manual Skoog Budget-Friendly Options
6. Navigating Solution Manual Skoog eBook Formats

- ePub, PDF, MOBI, and More
- Solution Manual Skoog Compatibility with Devices
- Solution Manual Skoog Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Manual Skoog
 - Highlighting and Note-Taking Solution Manual Skoog
 - Interactive Elements Solution Manual Skoog
- 8. Staying Engaged with Solution Manual Skoog
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Skoog
- 9. Balancing eBooks and Physical Books Solution Manual Skoog
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Skoog
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Manual Skoog
 - Setting Reading Goals Solution Manual Skoog
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution Manual Skoog
 - Fact-Checking eBook Content of Solution Manual Skoog
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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