Kame _____

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Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify the given (with units) and what you want to find (with units)
- c. Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Equation:	2 H ₂ O ₍₀₎ -	→ 2 H _{2(c)} +	1 O2 w
Before	2 mol	0 mool	0 mod
Change	-2 mod	+2 mol	
After	0 mol	2 mool	

$$36 \text{ g } H_1Ox \frac{1 \text{ mol } H_1O}{18.0 \text{ g } H_1O} = 2.0 \text{ mol } H_1O \times \frac{2 \text{ mol } H_2}{2 \text{ mol } H_1O} = 2 \text{ mol } H_2 \times \frac{2.0 \text{ g } H_2}{1 \text{ mol } H_2} = 4.0 \text{ g } H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

Equation:	2 Na ₍₀₎ +	_1_ Cl ₂₀₀ →	2 NaCloo	1100
Before	9.94 mol	xs mol	0 mol	
Change	-9.94 mod		+9.94 mol	
After	0 mol		9.94 mol	

$$581 \ g \ NaClx \\ \frac{1 \ mol \ NaCl}{58.5 \ g \ NaCl} = 9.93 \ mol \ NaCl \\ \times \\ \frac{2 \ mol \ Na}{2 \ mol \ NaCl} = 9.93 \ mol \ Na \\ \times \\ \frac{23.0 \ g \ Na}{1 \ mol \ Na} = 228 \ g \ NaCl} = 228 \ g \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 228 \ g \ NaCl} = 9.93 \ mol \ Na \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 228 \ g \ NaCl} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol \ NaCl \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 100 \ mol$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, C6H12O6, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?

Equation:	1 C6H12O600 +	6 O _{20(g)} →	6 CO _{2(g)}	+ 6 H ₂ O _(g)
Before	1.0 mol	as mol	0 mol	0 mol
Change	-1.0 mol	-1.0 mol	1000000	
After	0 med	as mod		

$$180.0 \text{ g C}_a \text{H}_{12} \text{O}_a \times \frac{1 \text{ mol C}_a \text{H}_{12} \text{O}_a}{180 \text{ g C}_a \text{H}_{12} \text{O}_a} = 1.00 \text{ mol C}_a \text{H}_{12} \text{O}_a \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_a \text{H}_{12} \text{O}_a} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol } O_2 \times \frac{32.0 \text{ g } O_2}{1 \text{ mol } O_2} = 192.0 \text{ g } O_2$$

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