

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L CO}_2 * \frac{(1 \text{ mol CO}_2)}{(22.4 \text{ L CO}_2)} = 3 \text{ mol CO}_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2 * \frac{(13 \text{ mol O}_2)}{(8 \text{ mol CO}_2)} = 4.875 \text{ mol O}_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol O}_2 \frac{(22.4 \text{ L O}_2)}{(1 \text{ mol O}_2)} = 109.2 \text{ L O}_2 = \boxed{109 \text{ L O}_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g Li} * \frac{(1 \text{ mol Li})}{(7 \text{ g Li})} = 2 \text{ mol Li}$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol Li} * \frac{(1 \text{ mol Li}_2\text{SO}_4)}{(2 \text{ mol Li})} = 1 \text{ mol Li}_2\text{SO}_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol Li}_2\text{SO}_4 * \frac{(110 \text{ g Li}_2\text{SO}_4)}{(1 \text{ mol Li}_2\text{SO}_4)} = \boxed{110 \text{ Li}_2\text{SO}_4}$$

Stoichiometry 2 Answers

Clifford Lane



Stoichiometry 2 Answers:

Foundations of Analytical Chemistry Miguel Valcárcel Cases,Ángela I. López-Lorente, Ma Ángeles

López-Jiménez,2017-08-29 This book offers a completely new approach to learning and teaching the fundamentals of analytical chemistry It summarizes 250 basic concepts of the field on the basis of slides Each of the nine chapters offers the following features Introduction Summary General scheme Teaching objectives Text containing the explanation of each slide Recommended and commented bibliography Questions to be answered Slides A distinct feature of this novel book is its focus on the fundamental concepts and essential principles of analytical chemistry which sets it apart from other books presenting descriptive overviews of methods and techniques *Microbiology (Questions and Answers), 5e* Purshotam Kaushik & Kirti

Kaushik,2022 Microbiology is an engaging textbook presenting balanced and comprehensive account of major areas of microbiology in the form of questions and answers This question answer approach to present complex topics and theories of microbiology regarding cellular and non cellular microorganisms microbial genetics and molecular biology in higher plants and animals makes the subject interesting and easily comprehensible for the students **STOICHIOMETRY** NARAYAN

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Stoichiometry and Thermodynamics of Metallurgical Processes
Y. K. Rao, 1985-10-31 Originally published in 1985 this textbook provides a thorough and comprehensive coverage of a wide range of topics in stoichiometry and thermodynamics with special emphasis on applications to metallurgical processes This book will be welcomed as a text for courses in elementary and advanced thermodynamics and stoichiometry

Electrochemical Science and Technology Keith Oldham, Jan Myland, Alan Bond, 2011-11-21 Electrochemistry is a discipline of wide scientific and technological interest Scientifically it explores the electrical properties of materials and especially the interfaces between different kinds of matter Technologically electrochemistry touches our lives in many ways that few fully appreciate for example materials as diverse as aluminum nylon and bleach are manufactured electrochemically while the batteries that power all manner of appliances vehicles and devices are the products of electrochemical research Other realms in which electrochemical science plays a crucial role include corrosion the disinfection of water neurophysiology sensors energy storage semiconductors the physics of thunderstorms biomedical analysis and so on This book treats electrochemistry as a science in its own right albeit resting firmly on foundations provided by chemistry physics and mathematics Early chapters discuss the electrical and chemical properties of materials from which electrochemical cells are constructed The behavior of such cells is addressed in later chapters with emphasis on the electrodes and the reactions that occur on their surfaces The role of transport to and from electrodes is a topic that commands attention because it crucially determines cell efficiency Final chapters deal with voltammetry the methodology used to investigate electrode behavior Interspersed among the more fundamental chapters are chapters devoted to applications of electrochemistry electrosynthesis power sources green electrochemistry and corrosion Electrochemical Science and Technology is addressed to all who have a need to come to grips with the fundamentals of electrochemistry and to learn about some of its applications

It will constitute a text for a senior undergraduate or graduate course in electrochemistry. It also serves as a source of material of interest to scientists and technologists in various fields throughout academia, industry, and government: chemists, physicists, engineers, environmentalists, materials scientists, biologists, and those in related endeavors. This book provides a background to electrochemistry as well as treating the topic itself. It is accessible to all with a foundation in physical science, not solely to chemists. It is addressed both to students and those later in their careers. Features web links through www.wiley.com/go/EST to extensive material that is of a more tangential, specialized, or mathematical nature. Includes questions as footnotes to support the reader's evolving comprehension of the material, with fully worked answers provided on the web. Provides web access to Excel spreadsheets which allow the reader to model electrochemical events. Has a copious Appendix of relevant data.

CBSE Chemistry Chapterwise Case Study Class 11 Priti Singhal, 2024-11-17. This book is structured to align with the latest syllabus and curriculum guidelines, ensuring that the content is both relevant and rigorous. Each chapter begins with a clear set of learning objectives, providing a roadmap for students to understand what they will achieve by the end of the chapter. We have included numerous diagrams, illustrations, and real-life examples to make complex concepts more accessible and engaging.

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