

Unit 1 - Stoichiometric Relationships

Terms

- A chemical element is a single pure substance, made of only one type of atom.
- A compound is a chemical combination of different elements, containing a fixed ratio of atoms.
 - Different physical properties
 - Different chemical properties
- A mixture is composed of two or more substances in which no chemical combination has occurred.
 - Homogenous mixtures have a uniform composition (evenly dispersed particles) (e.g. milk)
 - Heterogenous mixtures have a no uniformity (different dispersal) (e.g. cookie dough)
- Temperature is a measure of the average kinetic energy of the particles of a substance.
- Relative atomic mass, A_r , is the weighted average of the isotopes of one atom of an element relative to 1/12 of an atom of carbon-12.
- Relative formula mass, M_r , is the sum of the weighted average of the masses of the atoms in a formula unit relative to one-twelfth of an atom of carbon-12.
- The molar mass of a substance, M , is its relative atomic mass, A_r , or its relative formula mass, M_r , expressed in grams. It has the unit g mol^{-1} .
- One mole is the mass of substance that contains as many particles as there are atoms in 12 g of carbon-12.
- Limiting Reactant: Reactant that limits the amount of product that can be formed
- (e.g. $A + B \rightarrow C + D$; once A or B finishes, the reaction is over, and the one that finishes first is the limiting reactant and the other is in excess
- Theoretical Yield: the max amount of product that could be formed in a perfect scenario.
- Actual Yield: the amount of product formed experimentally

Important Formulae

- To find the number of moles in a solid, use
 - $n \text{ (moles)} = M \text{ (mass in grams)} / M_r \text{ (Molar Mass in grams)}$
- To calculate the number of particles (atoms, molecules, etc) in a sample, use
 - $n \text{ (moles)} \times N_A \text{ (avagadro's number: } 6.02 \times 10^{23})$
- To calculate the % yield and the efficiency of an exp. (must use the same unit), use
 - $(\text{actual amount} / \text{theoretical amount}) \times 100$
- To calculate the number of moles in a gas, use
 - $n \text{ (moles)} = \text{volume of gas (dm}^3) / \text{molar volume of the gas (dm}^3)$
- Ideal Gas Equation - remember the units in which each variable is calculated
 - $pV = nRT$ p (pressure in Pascals) $\times V$ (volume in m^3) = n (moles) $\times R$ (gas constant) $\times T$ (temperature in Kelvin)
- To calculate the number of moles in a solution, use
 - $n \text{ (moles)} = c \text{ (concentration in moles/dm}^3) \times v \text{ (volume in dm}^3)$

Stoichiometric Study Guide Answers

LM Reder



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Stoichiometric Study Guide Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Stoichiometric Study Guide Answers**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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