

CHAPTER

02

SOLUTION

1. **Solution** : A homogeneous mixture of two or more substances is known as solution.
2. **Solute** : The substance present in smaller amount in a solution is called solute.
3. **Solvent** : The substance present in larger amount in a solution is called solvent.
4. **Binary solution** : Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
5. **Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
6. **Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
7. **Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
8. **Weight Fraction** = $\frac{\text{weight of solute}}{\text{weight of solution}}$
9. **Weight percentage**

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
10. **Volume percentage**

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
11. **Volume Fraction** = $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
12. **Molarity (M)**

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1}{V(\text{lit})}$$
13. **Molarity** = $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml} \times \frac{1000}{1000}}$
14. **Molarity** = $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1000}{V(\text{ml})}$$
15. **Weight of solute** = $\frac{M \times GMW \times V \text{ in ml}}{1000}$
16. **Molarity for dilute solutions** : $M_1 V_1 = M_2 V_2$
17. **For neutralisation reactions** : $\frac{M_1 V_1}{n_1} = \frac{M_2 V_2}{n_2}$
18. **For mixed solution molarity**

$$= \frac{M_1 V_1 + M_2 V_2 + \dots}{V_1 + V_2 + \dots}$$

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Table of Contents Solutions Chapter 12 Review Study Guide Holt

1. Understanding the eBook Solutions Chapter 12 Review Study Guide Holt
 - The Rise of Digital Reading Solutions Chapter 12 Review Study Guide Holt
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Chapter 12 Review Study Guide Holt
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Chapter 12 Review Study Guide Holt
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Chapter 12 Review Study Guide Holt
 - Personalized Recommendations
 - Solutions Chapter 12 Review Study Guide Holt User Reviews and Ratings
 - Solutions Chapter 12 Review Study Guide Holt and Bestseller Lists
5. Accessing Solutions Chapter 12 Review Study Guide Holt Free and Paid eBooks
 - Solutions Chapter 12 Review Study Guide Holt Public Domain eBooks
 - Solutions Chapter 12 Review Study Guide Holt eBook Subscription Services
 - Solutions Chapter 12 Review Study Guide Holt Budget-Friendly Options
6. Navigating Solutions Chapter 12 Review Study Guide Holt eBook Formats

- ePub, PDF, MOBI, and More
 - Solutions Chapter 12 Review Study Guide Holt Compatibility with Devices
 - Solutions Chapter 12 Review Study Guide Holt Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions Chapter 12 Review Study Guide Holt
 - Highlighting and Note-Taking Solutions Chapter 12 Review Study Guide Holt
 - Interactive Elements Solutions Chapter 12 Review Study Guide Holt
 8. Staying Engaged with Solutions Chapter 12 Review Study Guide Holt
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Chapter 12 Review Study Guide Holt
 9. Balancing eBooks and Physical Books Solutions Chapter 12 Review Study Guide Holt
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Chapter 12 Review Study Guide Holt
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Solutions Chapter 12 Review Study Guide Holt
 - Setting Reading Goals Solutions Chapter 12 Review Study Guide Holt
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Solutions Chapter 12 Review Study Guide Holt
 - Fact-Checking eBook Content of Solutions Chapter 12 Review Study Guide Holt
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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