

Chapter 10 Study Guide - KEY

1. What is the molarity of a 2.05 L solution that is made by dissolving 27.5 g of sodium chloride in water?

$$\frac{27.5 \text{ g NaCl}}{58.5 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.470 \text{ mol}$$

$$M = \frac{\text{mol solute}}{\text{L sol'n}} = \frac{0.470 \text{ mol}}{2.05 \text{ L}} = \underline{0.229 M}$$

2. What is the normality of a 6.0 M solution of phosphoric acid?

$$N = (M)(\text{total pos. ox}) = (6.0 M)(3) = \underline{18 N}$$

3. What is the molality of a solution that is made by dissolving 98.7 g of calcium chloride in 755 g of water?

$$\frac{98.7 \text{ g CaCl}_2}{111.1 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.888 \text{ mol}$$

$$m = \frac{\text{mol solute}}{\text{kg solvent}} = \frac{0.888 \text{ mol}}{0.755 \text{ kg}} = \underline{1.18 m}$$

4. How many grams of KI are needed to prepare 325 mL of a 1.33 M solution?

$$1.33 M = \frac{n \text{ mol}}{0.325 \text{ L}}$$

$$n = 0.432 \text{ mol KI}$$

$$\text{mass} = \frac{0.432 \text{ mol KI}}{1 \text{ mol}} \times \frac{166.1 \text{ g}}{1} = \underline{71.8 \text{ g KI}}$$

5. How many grams of NaCl are needed to prepare a 4.15 m solution, using 176 g of water?

$$4.15 m = \frac{n \text{ mol}}{0.176 \text{ kg}}$$

$$n = 0.7304 \text{ mol NaCl}$$

$$\text{mass} = \frac{0.7304 \text{ mol NaCl}}{1 \text{ mol}} \times \frac{58.5 \text{ g}}{1} = \underline{42.7 \text{ g NaCl}}$$

6. Ethyl iodide ($\text{C}_2\text{H}_5\text{I}$) boils at 72.5°C and has a density of 1.933 g/mL .

- a. A solution prepared by dissolving 0.300 mol of a nonelectrolyte in 750.0 mL of ethyl iodide boils at 73.5°C . What is the boiling point constant (K_b) for ethyl iodide?

$$\text{C}_2\text{H}_5\text{I} = 156.0 \text{ g/mol}$$

$$\text{mass solvent} = \frac{750.0 \text{ mL}}{1 \text{ mL}} \times \frac{1.933 \text{ g}}{1} = 1449.75 \text{ g solvent}$$

$$m = \frac{0.300 \text{ mol}}{1.44975 \text{ kg}} = 0.207 m$$

$$\Delta T_b = K_b \cdot m$$

$$1.0^\circ\text{C} = (K_b)(0.207 m)$$

$$K_b = \underline{4.83^\circ\text{C/m}}$$

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The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

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