

Redox reaction

No. of equivalents, equivalent mass, normality

Equivalent concept $\rightarrow$ reaction independent
 $\rightarrow$ reaction doesn't need to be balanced

Reaction $\begin{cases} \rightarrow \text{Balanced} \rightarrow \text{use mole concept} \\ \rightarrow \text{Unbalanced} \begin{cases} \rightarrow \text{Balance it} \rightarrow \text{mole concept} \\ \rightarrow \text{Calculate } n_f \rightarrow \text{eq. concept} \end{cases} \end{cases}$

No. of equivalent $\rightarrow$ no. of moles $\times$ n-factor

$$\Sigma = \frac{\text{Molar mass}}{n_f}$$

$$\text{Normality} = \text{Molarity} \times n_f$$

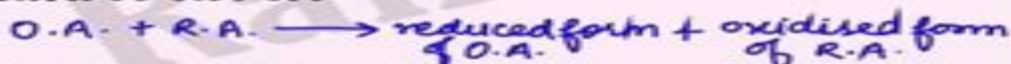
$$\Sigma = \frac{\text{Molar mass}}{\text{Total no. of } [H^+/\text{OH}^-] \text{ replaceable or to be replaced}}$$

is calculated for 1 mole only
 No. of e^- transferred per mole

$$N = \text{no. of equivalent/v}$$

Calculation of n-factor for different type of redox reaction-

$\rightarrow$ SIMPLE REDOX



$\rightarrow$ COMBINED REDOX



$\rightarrow$ SELF DECOMPOSITION REACTION



$$n_f = n \text{ Reduction}$$

In a single molecule, one element will undergo oxi^n & other one is reduced

$\rightarrow$ DISPROPORTIONATION REACTION



$$n_f = \frac{n_o \times n_R}{n_o + n_R}$$

same element gets oxidised as well as reduced



$$n = \frac{\text{total charge}}{\text{total item}} = \frac{2(7-2) + 1(7-4)}{3}$$

only for rea^n in which same element gets reduced in 2 diff. stages

$\rightarrow$ if oxidised $\rightarrow$ combined redox

Redox reaction

it is basically transfer of e^-



Law of equivalence

equivalent of acid = equivalent of base



Balancing of chemical reaction

- $\rightarrow$ Calculate n-factor of oxidising agent & reducing agent.
- $\rightarrow$ Cross multiply with the simplest ratio of n-f
- $\rightarrow$ Balance atoms under-going oxidation & reduction.
- $\rightarrow$ To Balance O & H

Acidic medium

O $\rightarrow$ comes from H_2O
 H $\rightarrow$ from H^+

Basic medium

O $\rightarrow$ comes from OH^-
 H $\rightarrow$ from H_2O



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Kenneth L. Henold



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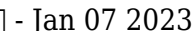
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