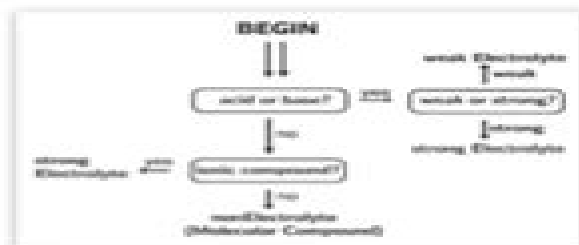


Aqueous Solutions

- Strong electrolytes fully dissolve and completely ionize. Strong electrolytes include salts like NaCl, strong acids like HCl, and strong bases like NaOH.
- For now just know that acids are chemicals that are able to donate a H^+ in a solution and a base is a chemical that is able to donate an OH^- in a solution. We will go over more details about these later.
- You should become familiar with the following strong acids and bases.

Strong Acids	Strong Bases
HCl	LiOH $\text{Ca}(\text{OH})_2$
HBr	NaOH $\text{Sr}(\text{OH})_2$
HI	KOH
HClO_3	RbOH
HClO_4	CsOH
HNO_3	
H_2SO_4	

- A weak electrolyte is a weak acid or weak base. Weak electrolyte only partially ionize in solution. If the acid or base is not part of the list you should familiar with then assume it is a weak acid or weak base.
- Non-Electrolytes are covalent compounds. However, that does not mean the chemical is not soluble. Sugar is very soluble in water, but it does not ionize in solution because it is made of Carbon, Hydrogen, and Oxygen all which are nonmetals thus a covalent bond.
- You can use the flow chart to determine if a chemical is a strong, weak, or non-electrolyte.



Molarity and Dilutions

- If we take any chemical and put it on a scale we can find the number of moles by the molar mass and this will tell us how many atoms or molecules that we have of that

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