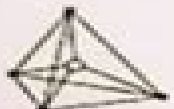


Summary of Molecular Shapes

Type of Molecule	Molecular Shape	Atoms Bonded to Central Atom	Lone Pairs of Electrons	Formula Example	Lewis Structure
Linear		2	0	BeF_2	$\text{:}\ddot{\text{F}}\text{—Be—}\ddot{\text{F}}\text{:}$
Bent		2	1	SnCl_2	$\text{:}\ddot{\text{Cl}}\text{—}\ddot{\text{Sn}}\text{—}\ddot{\text{Cl}}\text{:}$
Trigonal-planar		3	0	BF_3	$\text{:}\ddot{\text{F}}\text{—}\ddot{\text{B}}\text{—}\ddot{\text{F}}\text{:}$ $\text{:}\ddot{\text{F}}\text{:}$
Tetrahedral		4	0	CH_4	$\begin{array}{c} \text{H} \\ \\ \text{H—C—H} \\ \\ \text{H} \end{array}$
Trigonal-pyramidal		3	1	NH_3	$\begin{array}{c} \text{:}\ddot{\text{N}}\text{:} \\ \\ \text{H—H—H} \end{array}$
Bent		2	2	H_2O	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\ \\ \text{H—H} \end{array}$
Trigonal-bipyramidal		5	0	PCl_5	$\begin{array}{c} \text{:}\ddot{\text{Cl}}\text{:} \quad \text{:}\ddot{\text{Cl}}\text{:} \\ \quad \quad \\ \text{:}\ddot{\text{Cl}}\text{—}\ddot{\text{P}}\text{—}\ddot{\text{Cl}}\text{:} \\ \quad \quad \\ \text{:}\ddot{\text{Cl}}\text{:} \quad \text{:}\ddot{\text{Cl}}\text{:} \end{array}$
Octahedral		6	0	SF_6	$\begin{array}{c} \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \\ \quad \quad \quad \quad \\ \text{:}\ddot{\text{F}}\text{—}\ddot{\text{S}}\text{—}\ddot{\text{F}}\text{—}\ddot{\text{F}}\text{:} \\ \quad \quad \quad \quad \\ \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \quad \text{:}\ddot{\text{F}}\text{:} \end{array}$

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