

THE CULTURE OF AI

everyday life and the digital revolution

ANTHONY ELLIOTT



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web problem 6 1 6 1 6 calculate the molarity of each of the following solutions a 0 195 g of cholesterol $\text{C}_{27}\text{H}_{46}\text{O}$ in 0 100 l of serum the average concentration of cholesterol in human serum b 4 25 g of NH_3 in 0 500 l of solution the concentration of NH_3 in household ammonia

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web 1 molarity $M = \frac{\text{mass molar mass}}{x}$ 0 4500 l 0 825 g 141 9579 g mol^{-1} x 0 0129 m 2 molality 0 825 g 141 9579 g mol^{-1} 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction n_2 $\frac{\text{hpo}_4}{0 825 \text{ g } 141 9579 \text{ g } \text{mol}^{-1} 0 00581158 \text{ mol } \text{H}_2\text{O}}$ 450 0 g 18 015 g mol^{-1} 24 97918401 mol

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based

on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it s written like this
$$\frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100$$
 end array

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web aug 13 2017 molarity practice problems this chemistry video tutorial explains how to solve common molarity problems it discusses how to calculate the concentration of a solution given the mass in grams