

Solution and Answer Guide

Solution and Answer Guide: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45039-0, Chapter 9: Solving Equilibrium Problems for Complex Systems

Chapter 9

- 9-1.** Demonstrate how the sulfide ion concentration is related to the hydronium ion concentration of a solution that is kept saturated with hydrogen sulfide.

Solution:

The overall dissociation constant for H_2S is

$$\frac{[\text{H}_3\text{O}^+]^2[\text{S}^{2-}]}{[\text{H}_2\text{S}]} = K_1K_2$$

In a solution saturated with the gas, $[\text{H}_2\text{S}]$ is constant, therefore

$$[\text{S}^{2-}] = \frac{[\text{H}_2\text{S}]}{[\text{H}_3\text{O}^+]^2} K_1K_2 \text{ or } [\text{S}^{2-}] = \frac{K}{[\text{H}_3\text{O}^+]^2}$$

- *9-2.** Why are simplifying assumptions in equilibrium problems restricted to relationships that are sums or differences?

Answer:

To simplify equilibrium calculations, we sometimes assume that the concentrations of one or more species are negligible and can be approximated as 0.00 M. In a sum or difference assuming a concentration is 0.00 M leads to an appropriate result. In contrast, if we were to simplify an equilibrium constant expression by assuming one or more concentrations are zero, we would be multiplying or dividing by 0.00, which would render the expression meaningless.

- 9-3.** In the margin note on page 215, we suggest that the term *mass-balance equation* may be something of a misnomer. Using a specific chemical system, discuss mass balance and show that mass balance and concentration balance are equivalent.

Answer:

Let us take the case of 0.100 M acetic acid (HOAc). We usually write the mass balance as a concentration balance. That is

$$C_{\text{HOAc}} = 0.100 = [\text{HOAc}] + [\text{OAc}^-]$$

Since the concentrations are moles/L, we can write

$$\text{Initial moles HOAc/L} = \text{equilibrium moles HOAc/L} + \text{equilibrium moles OAc}^-/\text{L}$$

*Answers and solutions are provided in the Student Solutions Manual for questions and problems marked with an asterisk.

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**Alan Langford, John R. Dean, Robert H.
Reed, Jonathan D. B. Weyers, David
Holmes, Allan M. Jones**

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Table of Contents Skoog Spectroscopy Answer Guide

1. Understanding the eBook Skoog Spectroscopy Answer Guide
 - The Rise of Digital Reading Skoog Spectroscopy Answer Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Skoog Spectroscopy Answer Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skoog Spectroscopy Answer Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skoog Spectroscopy Answer Guide
 - Personalized Recommendations
 - Skoog Spectroscopy Answer Guide User Reviews and Ratings
 - Skoog Spectroscopy Answer Guide and Bestseller Lists
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 - Interactive Elements Skoog Spectroscopy Answer Guide
- 8. Staying Engaged with Skoog Spectroscopy Answer Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skoog Spectroscopy Answer Guide
- 9. Balancing eBooks and Physical Books Skoog Spectroscopy Answer Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skoog Spectroscopy Answer Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skoog Spectroscopy Answer Guide
 - Setting Reading Goals Skoog Spectroscopy Answer Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skoog Spectroscopy Answer Guide
 - Fact-Checking eBook Content of Skoog Spectroscopy Answer Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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