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INVESTING SIMPLIFIED

BUILDING WEALTH ONE STEP AT A TIME

Investing Simplified Ebook

Matthew Partridge



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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kJ mol⁻¹

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eg 1 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o l}$ for which $\Delta H = 57\text{ kJ mol}^{-1}$ eg 2 $\text{h}_2\text{so}_4 \text{ aq} + \text{koh aq} \rightarrow \text{ks}_2\text{o}_7 \text{ aq} + \text{h}_2\text{o l}$ for which $\Delta H = 57\text{ kJ mol}^{-1}$

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