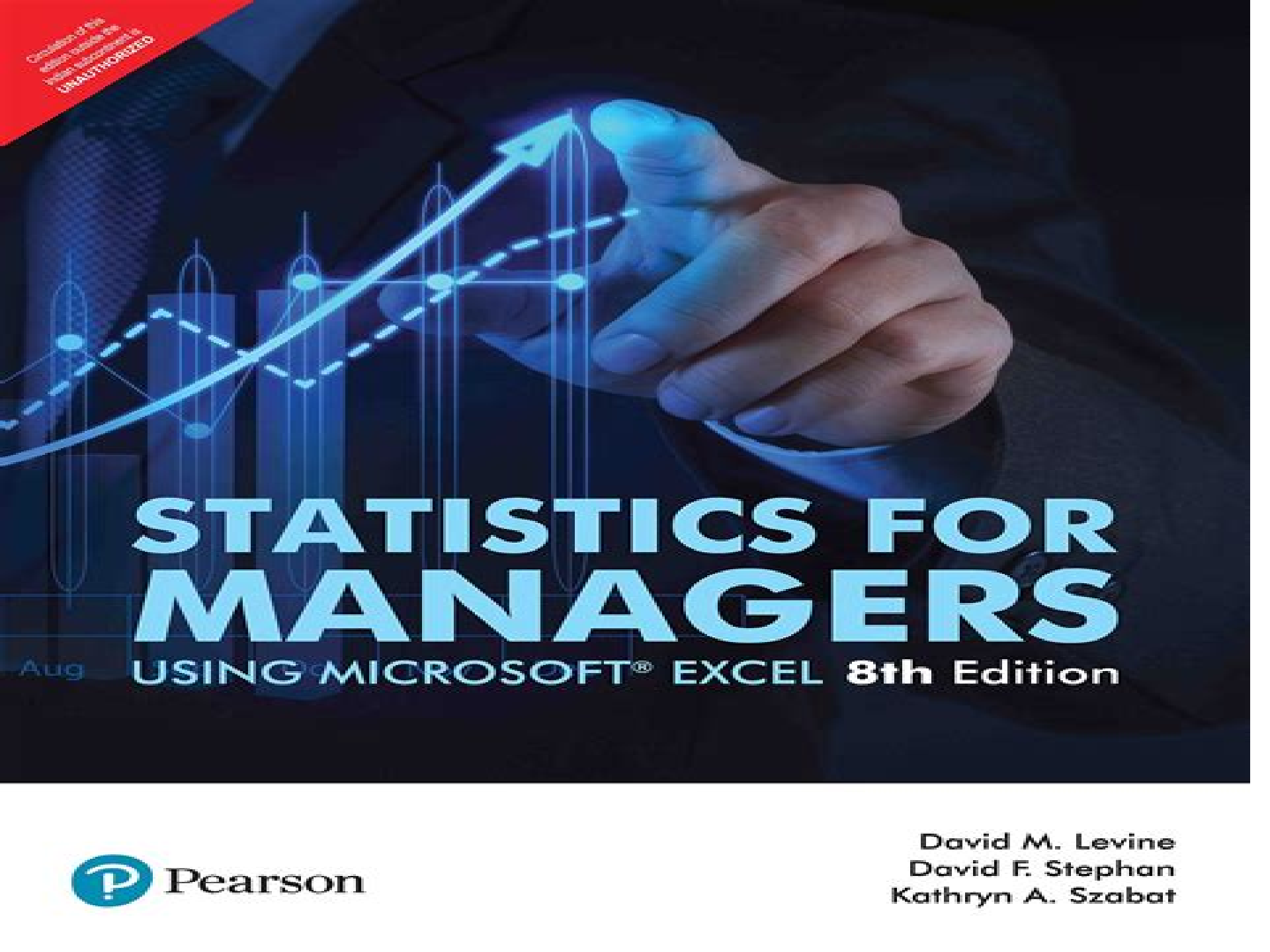


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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{Ag}^+ \text{aq} + 2 \text{F}^- \text{aq} + 2 \text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{NH}_4^+ \text{aq} + 2 \text{F}^- \text{aq}$ because two NH_4^+ aq

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