

Answers to Chapter 3 Study Questions

- $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$ b) 92.1 g c) $6.02 \times 10^{23} \text{ molecules}$
 - $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
- 4 atoms (one N + 3 H)
 - $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$
 - $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
- Molar mass of $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$; $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$; $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$
33.3% Na, 20.3% N and 46.4% O.
- In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$ $3.39/3.39 = 1 \times 2 = 2$
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$ $5.1/3.39 = 1.5 \times 2 = 3$
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$ $3.39/3.39 = 1 \times 2 = 2$
empirical formula = $\text{C}_2\text{H}_3\text{O}_2$
 - Molar mass of $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$
 $118/59.0 = 2 \rightarrow$ molecular formula = **$\text{C}_4\text{H}_6\text{O}_4$**
- In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$.
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$ $0.296/0.296 = 1$
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$ $0.296/0.296 = 1$
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$ $0.891/0.296 = 3$
formula = MgCO_3
 - $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$; $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$; $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$. (You should get the same result using molar mass and atomic masses.)
 - $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

Stars Suite Chemistry Answers

Justin P. Lomont, Ian C. Stewart



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The Handy Chemistry Answer Book Justin P. Lomont, Ian C. Stewart, 2013-10-01 Don't be mixed up about chemistry. Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy to follow question and answer guide. Where would we be without atoms and compounds? Gas, liquids, solids, and plasma. Acids and bases. Bonds and reactions. Matter and energy. The Handy Chemistry Answer Book covers the building blocks of life and the universe. The secret life of atoms, how polar bears aren't actually white, why oil and water don't mix, and much, much more are revealed and explained. This informative guide covers the basics of chemistry: history, atomic structures, chemical bonds, and reactions; organic and inorganic chemistry; to more advanced material: nuclear chemistry, biochemistry, physical and theoretical chemistry. By answering nearly 1,000 common chemistry questions including: What causes lightning? How does photosynthesis work? What are hard and soft Lewis acids and bases? What makes a fabric waterproof? What are the twelve principles of green chemistry? When did alchemists finally abandon trying to make gold? What is Le Chatelier's principle? What do the different octane ratings mean at the gas pump? What is genetic engineering? Why is calcium important for strong bones? What is the 18 electron rule? Why does chocolate turn white as it ages? Chemical reactions that rule the world: their properties, structure, composition, behavior, and history are tackled and explained in plain English. In *The Handy Chemistry Answer Book*, with many photos, illustrations, a few formulas, molecular diagrams, and other graphics, this fun fact filled tome is richly illustrated. A history of chemistry, timeline, appendices on Nobel Prize winners, a bibliography, further reading section, glossary of terms, a table of physical constants, a table of conversion factors, and extensive index add to its usefulness. *Chemical Engineering Abstracts*, 1986. **Five Star Fraud** Margaret G. Cahill, 2010-02-04 Neil and Bridget Landers relocate to Charlotte, North Carolina for work. Quickly, Neil learns the affable, generous CEO that recruited him is vindictive and capriciously assigns blame and punishes employees for minor infractions. Neil keeps a low profile until he clashes with the boss son over exaggerated marketing claims. Reluctantly, Neil blows the whistle. This triggers a battle to save his reputation, his marriage, and his life. P. 4 of cover. **The Star Shiner** Evan Richardson, 2013-05-11 The Star Shiner poses the question: Can a young man from a small rural Kentucky town, fleeing a domineering mother and an abusive alcoholic father, find recognition and happiness in New York City working with the high-powered fashion and cosmetic industries and with some of the world's most famous people without losing his values and his soul? After a notable career in fashion illustration and modeling in Paris, Richardson becomes a makeup artist working with *Vogue* and *Harper's Bazaar* and photographers Richard Avedon, Irving Penn, and Francesco Scavullo. The book is more than a celebrity memoir; in that it is particularly a narrative of the 70s into the 90s, a time in New York's history of financial difficulties and corruption. This was the golden era of fashion and cosmetics when *Vogue* editor in chief Diana Vreeland, Charles Revson, and Estée Lauder, the king and queen of cosmetics, and Avedon, Penn, and Scavullo ruled the industries. The reader is taken into wild nights at Studio 54 and into New York's

downtown after hours dungeons seething with sex drugs and danger where individuals are treated as sexual stepping stones
Then the party ended with the scourge of AIDS that took Richardson s partner and turned Richardson s life around with a
spiritual awakening **Boston Journal of Chemistry and Pharmacy** ,1866 *Job Safety & Health Quarterly* ,1994

Departments of Veterans Affairs and Housing and Urban Development, and Independent Agencies Appropriations for
Fiscal Year 2002 United States. Congress. Senate. Committee on Appropriations. Subcommittee on VA-HUD-Independent
Agencies,2002 **Marrying the Lone Star Maverick** Sara Orwig,2010-01-21 He needed a wife fast And Texan Jeff Brand s
lovely new assistant would do just fine After all the heat between him and Holly Lombard was becoming impossible to resist
And a no strings marriage would certainly work for them both Yet once he put the ring on her finger Jeff realized he may
have gotten in over his head He could ride a wild bronco run a multimillion dollar business and charm any woman he chose
But keeping feelings out of this in name only union With a wife like Holly the maverick groom had just discovered his
toughest challenge yet *Interior, Environment, and Related Agencies Appropriations for 2016: Justification of the budget
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