

001

REKI KAWAHARA

ILLUSTRATION BY abec



The illustration depicts two anime-style characters in a dynamic pose. On the left, a male character with short black hair and a serious expression is shown from the chest up, wearing dark, detailed armor. He holds a sword with a black hilt and a silver blade. On the right, a female character with long, flowing brown hair and a gentle smile is shown from the waist up, wearing a white and red outfit. She holds a sword with a silver hilt and a silver blade. The background is a dark, swirling mass of green and white, suggesting a forest or a digital environment. The title 'SWORD ART ONLINE' is prominently displayed in the center, with 'Aincrad' written below it in a stylized font. The overall style is characteristic of Japanese anime art.

SWORD ART ONLINE Aincrad

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Sword Art Online 1 Aincrad

Anthony S. Fauci



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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 $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{C}_6\text{H}_4\text{COONa} + \text{H}_2\text{O}$
 the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

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