

Stoichiometry, Limiting Reactants, Percent Yield, and Excess Reactant Remaining Worksheet

1. **Aluminum** reacts with **oxygen** to produce **aluminum oxide** which can be used as an adsorbent, desiccant or catalyst for organic reactions.

1 mole of aluminum (Al) is 26.982 g
1 mole of molecular oxygen (O₂) is 31.998 g
1 mole of aluminum oxide is 101.961 g

- a. What is the formula of aluminum oxide? (First hint: use the criss-cross rule with aluminum ion and oxide ion) (Second hint: check your answer with the molar mass of aluminum oxide)



- b. Write the balanced chemical equation for the reaction of aluminum (Al) with molecular oxygen (O₂) to make aluminum oxide.



- c. How many grams of aluminum oxide could form if 82.49 g Al reacts completely?

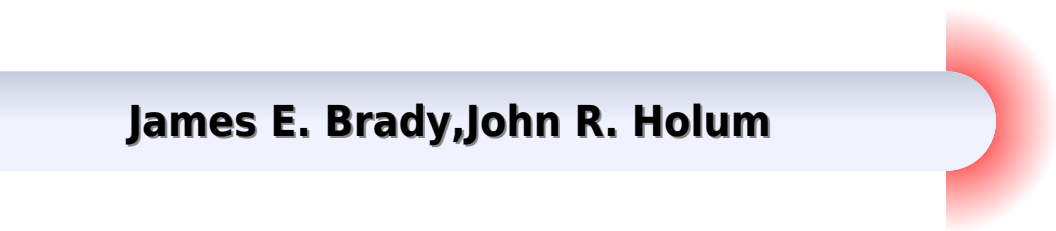
$$82.49 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{101.961 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 155.9 \text{ g Al}_2\text{O}_3$$

- d. How many grams of aluminum oxide could form if 117.65 g O₂ reacts completely?

$$117.65 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} \times \frac{101.961 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 249.9 \text{ g Al}_2\text{O}_3$$

Study Guide Stoichiometry Limiting Reactant Percent Yield

James E. Brady, John R. Holum



Study Guide Stoichiometry Limiting Reactant Percent Yield:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Study Guide to Accompany Calculus for the Management, Life, and Social Sciences* Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18-chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple-choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas, and the mole, chemical reactions, calculations based on equations, gases, and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics, and equilibrium, and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students. *The Complete Chemistry Study Guide and Note Cards and MCAT* Konstantinos Papadopoulos, 2012-07-06

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Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover

Know Your 'O' Level Chemistry - A Study Guide, Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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