

LAB: Specific Heat of Solids

Objective

The purpose of this lab was to determine the specific heat of three metals by only using the principles of calorimetry and fundamentals of thermodynamics.

Materials

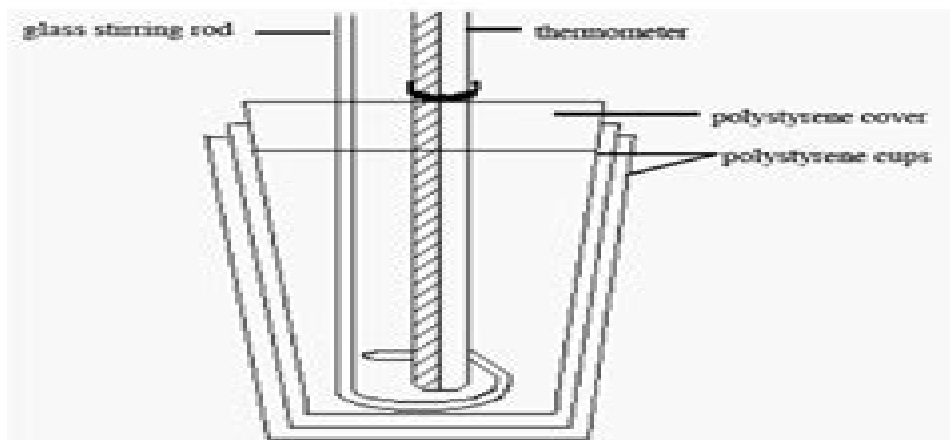
- ✓ Styrofoam coffee cups (2 per group) with cover
- ✓ Hot plates
- ✓ Graduated cylinders (200 mL)
- ✓ Glass thermometer
- ✓ String
- ✓ Metal samples -- a different metal sample for each group

Theory

As three different samples of metals are being heated to high temperatures they will be then placed in a cup containing a known amount of water. determine how much heat was gained by the water in the calorimeter to know how much heat was lost by the metals. calculate and compare results to known values.

Procedure

weight empty styrofoam cups with cover and record data. then measure 100.00mL of water using a graduated cylinder and pour water into the styrofoam cups. Now get the new mass and record data. Make sure to also record the initial temperature of water on the data table. Set up a hot water bath and set water to boil. weight the metal samples and tie a piece of string to them so it can be suspended in the water. Take the temperature of boiling water and record data. Once that is done, quickly take out the hot metal from boiling water to be placed in the water of the styrofoam cups and quickly cover so that the heat doesn't escape. gently agitate the styrofoam cups and record the final temperature of the water. Calculate the heat gained by the water, Q_w . Record this on your data Sheet. Using this number calculate the heat lost by the metal, Q_m . Now compare the experimental value with the actual value, using table 19-



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