

Specific Heat Capacity Lab Practice

Specific Heat Capacity_{water} = $4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$ $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$

A 25.00 g piece of metal at 110.0°C is placed in a calorimeter containing 15.00 g of water at 25.00°C . The final temperature of the system is 40.70°C . Determine the specific heat capacity of the metal.

Heat gained by water = $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$ or $c = \frac{15.00 \text{ g} \cdot 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \cdot (40.70^\circ\text{C} - 25.00^\circ\text{C})}{15.00 \text{ g}}$

$$\text{Heat lost by metal} = \text{Heat gained by water} \quad q = 25.00 \text{ g} \cdot c \cdot (110.0^\circ\text{C} - 40.70^\circ\text{C}) = 15.00 \text{ g} \cdot 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \cdot (40.70^\circ\text{C} - 25.00^\circ\text{C})$$

$$\frac{q}{m \cdot \Delta T} = \frac{15.00 \text{ g} \cdot 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}}{25.00 \text{ g} \cdot (110.0^\circ\text{C} - 40.70^\circ\text{C})} = c = \frac{15.00 \text{ g} \cdot 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \cdot (40.70^\circ\text{C} - 25.00^\circ\text{C})}{25.00 \text{ g} \cdot (110.0^\circ\text{C} - 40.70^\circ\text{C})}$$

$$c = 0.2583 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \quad \text{Answer}$$

Specific Heat Capacity_{water} = $4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$

A 40.0 Gg piece of metal at 100.0°C is placed in a calorimeter containing 30.00 g of water at 20.00°C . The final temperature of the system is 25.5°C . Determine the specific heat capacity of the metal.

Heat gained by H₂O = Heat lost by metal

Heat gained by H₂O = $q = m \cdot c \cdot \Delta T$ or $c = \frac{q}{m \cdot \Delta T}$ or $c = \frac{30.00 \text{ g} \cdot 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \cdot (25.5^\circ\text{C} - 20.00^\circ\text{C})}{30.00 \text{ g}}$

$$c = \frac{q}{m \cdot \Delta T} = \frac{15.00 \text{ J}}{15.00 \text{ g} \cdot (110.0^\circ\text{C} - 40.70^\circ\text{C})}$$

$$c = 0.2583 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}} \quad \text{Answer}$$

Specific Heat Capacity Lab Manual Answers

Judd E. Hollander

A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the author's name.

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