

CONCENTRATION OF SOLUTIONS

The concentration of a solution is usually measured in moles per cubic decimene (molitim³). This is a measure of the number of moles in one cubic deciments.

The volume must be in dry² (there are 1000 cm² in 1 dm³). Vol in dry³ = <u>vol in cm³</u>

concertisation (resister): malama danéh



- Calculate the concentration of the following solutions in recition².
 - at 10,110 motes of NaCli in 200 cm³
 - 50 . 0.20 motes of H-SO₄ in 100 cm²
 - ct. It 020 moins of NaCH is 25 cm2
- Calculate the number of moles in the following solutions.
 - at 100 cm³ of 0.20 molitim³ HHVDs
 - to 25 cm2 at 1.50 masses 160H.
 - ct. 50 cm² ef 0.10 merom² HuSQ.

Concentration can also be measured in grams per cubic deciration (gidna). This is a nemember that mass - At a

A simple conversion is:

Competition + 2 months **阿利尼岛岛**

Competition v 2 x 56 x 56 x 66 x 660

conc (gidim¹) = conc (molidm¹) x M.

- Calculate the concentration of the following solutions in grow³.
 - at: 0.100 molecus NaCH
 - b) 0.250 molitan³ DHyCOOH
 - di 150 matem¹ HNCs.
- 0.20 moles of NaCini is dissolved in 250 cm³ of water.
 - at Calculate the omcentration is molitimal
 - b). Calculate the concentration in older
- 5.0 g of KD4Ds is dissolved in 100 cm² of water.
 - a). Calculate the concentration in gittin³.
 - to Calculate the concentration in molitimal

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David R. Klein

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