

Solubility Rules Worksheet

1. Classify each of the substances as being soluble or insoluble in water.

- a. KBr =
- b. PbCO_3 =
- c. BaSO_4 =
- d. zinc hydroxide =
- e. sodium acetate =
- f. silver iodide =
- g. cadmium (II) sulfide =
- h. zinc carbonate =

- i. silver acetate =
- j. copper (II) sulfide =
- k. $\text{Mg}_3(\text{PO}_4)_2$ =
- l. KOH =
- m. NaCl =
- n. NH_4OH =
- o. Hg_2SO_4 =
- p. PbI_2 =

2. Identify the two new compounds which form if the solutions, as suggested by the following table, were mixed. CIRCLE the names of the compounds which would precipitate from the solutions.

	KBr	Na_2CO_3	CaS	NH_4OH
AgNO_3				
BaCl_2				
$\text{Al}(\text{NO}_3)_3$				
CuSO_4				

Answers to Solubility Rules Worksheet

3. Classify each of the substances as being soluble or insoluble in water.

- a. potassium bromide - **sol**
- b. lead (II) carbonate - **insol**
- c. barium sulfate - **insol**
- d. zinc hydroxide - **insol**
- e. sodium acetate - **sol**
- f. silver iodide - **insol**
- g. cadmium (II) sulfide - **insol**
- h. zinc carbonate - **insol**

- i. silver acetate - **sol**
- j. copper (II) sulfide - **insol**
- k. $\text{Mg}_3(\text{PO}_4)_2$ - **insol**
- l. KOH - **sol**
- m. NaCl - **sol**
- n. NH_4OH - **sol**
- o. Hg_2SO_4 - **insol**
- p. PbI_2 - **insol**

4. Identify the two new compounds which form if the solutions, as suggested by the following table, were mixed. CIRCLE the names of the compounds which would precipitate from the solutions.

	KBr	Na_2CO_3	CaS	NH_4OH
AgNO_3	<u>AgBr</u> + KNO_3	NaNO_3 + <u>Ag_2CO_3</u>	$\text{Ca}(\text{NO}_3)_2$ + <u>Ag_2S</u>	<u>AgOH</u> + NH_4NO_3
BaCl_2	KCl + BaBr_2	NaCl + <u>BaCO_3</u>	CaCl_2 + BaS	$\text{Ba}(\text{OH})_2$ + <u>NH_4Cl</u>
$\text{Al}(\text{NO}_3)_3$	AlBr_3 + KNO_3	NaNO_3 + <u>$\text{Al}_2(\text{CO}_3)_3$</u>	$\text{Ca}(\text{NO}_3)_2$ + <u>Al_2S_3</u>	NH_4NO_3 + <u>$\text{Al}(\text{OH})_3$</u>

Solubility Rules Lab Answer Key

Karen C. Timberlake



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