#### WORKSHEET ON SOLUTION CONCENTRATIONS

#### INTRODUCTION

A solution is made up of two parts, the **SOLVENT** and the **SOLUTE**. The solution could be **CONCENTRATED** or **DILUTE**.

SOLVENT	The liquid which dissolves the solute. It is usually water.		
SOLUTE	The substance which dissolves - often a solid, but not always.		
CONCENTRATED	There is a lot of solute and a little water.		
DILUTE	There is a lot of water and only a little solute.		

When we measure out 50cm<sup>3</sup> of a solution, e.g., hydrochloric acid, we measure out some solvent (water) and some acid together. In any reactions of the acid, it is the HCl particles which are involved. The water is just a carrier for the acid, so when we measure out a volume of the solution, we want to know how much acid it contains. To do this, we need to define the CONCENTRATION of the solution. This tells us how much solute there is in a fixed volume of the solution. In chemistry, "how much" means MOLES and the "fixed volume" is taken as 1 cubic decimetre (1dm<sup>3</sup>).

[NB: Idm<sup>3</sup> is 1000cm<sup>3</sup> and is the same volume as 1 litre.]

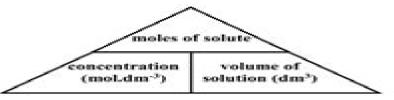
#### DEFINITION

Concentration of solution = (moles per dm³)		number of moles of solute
		volume of solution in dm <sup>3</sup>

- You must learn this relationship and be able to use it and rearrange it to calculate moles
  of solute or volume of solution, given the other two quantities.
- Remember that the volume must be in dm<sup>3</sup>, so convert from cm<sup>3</sup> if necessary.
- The units for concentration are moles per cubic decimetre which is written mol.dm<sup>-3</sup>
- Lab bottles use the letter M (meaning molar) and this is the same thing as mol.dm<sup>3</sup>

#### RELATIONSHIPS

Use the triangle opposite to help you answer the questions which follow. Some worked examples are given to help you.



#### Examples:-

- I) What is the concentration of a solution which contains 2 moles of copper (II) sulphate in 4dm³ of solution?
  - Ans: concentration = moles + volume =  $2 + 4 = 0.5 \text{ mol.dm}^3$
- 2) How many moles of sodium hydroxide are there in 200cm<sup>3</sup> of a solution of concentration 0.2 mol.dm<sup>-3</sup>? (Remember that 200cm<sup>3</sup> = 200 + 1000 = 0.2dm<sup>3</sup>)
  Ans: moles = concentration × volume = 0.2 × 0.2 = 0.04 moles

# **Skill Practice 39 Concentration Practice Answers**

**Kelly Hackett** 

#### **Skill Practice 39 Concentration Practice Answers:**

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#### Table of Contents Skill Practice 39 Concentration Practice Answers

- 1. Understanding the eBook Skill Practice 39 Concentration Practice Answers
  - The Rise of Digital Reading Skill Practice 39 Concentration Practice Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Skill Practice 39 Concentration Practice Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Skill Practice 39 Concentration Practice Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Skill Practice 39 Concentration Practice Answers
  - Personalized Recommendations
  - Skill Practice 39 Concentration Practice Answers User Reviews and Ratings
  - Skill Practice 39 Concentration Practice Answers and Bestseller Lists
- 5. Accessing Skill Practice 39 Concentration Practice Answers Free and Paid eBooks
  - Skill Practice 39 Concentration Practice Answers Public Domain eBooks
  - Skill Practice 39 Concentration Practice Answers eBook Subscription Services
  - Skill Practice 39 Concentration Practice Answers Budget-Friendly Options
- 6. Navigating Skill Practice 39 Concentration Practice Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Skill Practice 39 Concentration Practice Answers Compatibility with Devices
  - Skill Practice 39 Concentration Practice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Skill Practice 39 Concentration Practice Answers
  - Highlighting and Note-Taking Skill Practice 39 Concentration Practice Answers
  - Interactive Elements Skill Practice 39 Concentration Practice Answers

- 8. Staying Engaged with Skill Practice 39 Concentration Practice Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Skill Practice 39 Concentration Practice Answers
- 9. Balancing eBooks and Physical Books Skill Practice 39 Concentration Practice Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Skill Practice 39 Concentration Practice Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 39 Concentration Practice Answers
  - Setting Reading Goals Skill Practice 39 Concentration Practice Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 39 Concentration Practice Answers
  - Fact-Checking eBook Content of Skill Practice 39 Concentration Practice Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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