

WORKSHEET ON SOLUTION CONCENTRATIONS

INTRODUCTION

A solution is made up of two parts, the **SOLVENT** and the **SOLUTE**. The solution could be **CONCENTRATED** or **DILUTE**.

SOLVENT

The liquid which dissolves the solute. It is usually water.

SOLUTE

The substance which dissolves - often a solid, but not always.

CONCENTRATED

There is a lot of solute and a little water.

DILUTE

There is a lot of water and only a little solute.

When we measure out 50cm^3 of a solution, e.g., hydrochloric acid, we measure out some solvent (water) and some acid together. In any reactions of the acid, it is the HCl particles which are involved. The water is just a carrier for the acid, so when we measure out a volume of the solution, we want to know how much acid it contains. To do this, we need to define the **CONCENTRATION** of the solution. This tells us *how much* solute there is in a *fixed volume* of the solution. In chemistry, "*how much*" means **MOLES** and the "*fixed volume*" is taken as 1 cubic decimetre (1dm^3).

[NB: 1dm^3 is 1000cm^3 and is the same volume as 1 litre.]

DEFINITION

$\text{Concentration of solution (moles per dm}^3\text{)} = \frac{\text{number of moles of solute}}{\text{volume of solution in dm}^3}$

- You must learn this relationship and be able to use it and rearrange it to calculate moles of solute or volume of solution, given the other two quantities.
- Remember that the volume must be in dm^3 , so convert from cm^3 if necessary.
- The units for concentration are *moles per cubic decimetre* which is written mol.dm^{-3}
- Lab bottles use the letter *M* (meaning *molar*) and this is the same thing as mol.dm^{-3}

RELATIONSHIPS

Use the triangle opposite to help you answer the questions which follow. Some worked examples are given to help you.



Examples:-

- 1) What is the concentration of a solution which contains 2 moles of copper (II) sulphate in 4dm^3 of solution?

Ans: concentration = moles ÷ volume = $2 \div 4 = \underline{0.5 \text{ mol.dm}^{-3}}$

- 2) How many moles of sodium hydroxide are there in 200cm^3 of a solution of concentration 0.2 mol.dm^{-3} ? (Remember that $200\text{cm}^3 = 200 \div 1000 = 0.2\text{dm}^3$)

Ans: moles = concentration × volume = $0.2 \times 0.2 = \underline{0.04 \text{ moles}}$

Skill Practice 39 Concentration Practice Answers

Kelly Hackett



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