

Stoichiometry

Section 11.1 Defining Stoichiometry

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Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule N_2 + 3 molecules

$\text{H}_2 \rightarrow$ 2 molecules NH_3

1 mole N_2 + 3 moles $\text{H}_2 \rightarrow$ 2 moles NH_3

Mass:

N_2 : (2 mol)(14.007 g/mol) = 28.014 g

3H_2 : (6 mol)(1.008 g/mol) = 6.048 g

2NH_3 : (2 mol)(14.007 g/mol) +
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g N_2 + 6.048 g $\text{H}_2 \rightarrow$ 34.062 g NH_3

34.062 g reactants = 34.062 g products



1 molecule HCl + 1 formula unit $\text{KOH} \rightarrow$

1 formula unit KCl + 1 molecule H_2O

1 mole HCl + 1 mole $\text{KOH} \rightarrow$

1 mole KCl + 1 mole H_2O

Mass:

HCl : (1 mol)(1.008 g/mol) +
(1 mol)(35.453 g/mol) = 36.461 g

KOH : (1 mol)(39.098 g/mol) +
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)
= 56.105 g

KCl : (1 mol)(39.098 g/mol) +
(1 mol)(35.453 g/mol) = 74.551 g

H_2O : (2 mol)(1.008 g/mol) +
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g $\text{KOH} \rightarrow$
74.551 g KCl + 18.015 g H_2O

92.566 g reactants = 92.566 g products



2 atoms Mg + 1 molecule

$\text{O}_2 \rightarrow$ 2 formula units MgO

2 moles Mg + 1 mole $\text{O}_2 \rightarrow$ 2 moles MgO

Mass:

2Mg : (2 mol)(24.305 g/mol) = 48.610 g

O_2 : (2 mol)(15.999 g/mol) = 31.998 g

2MgO : (2 mol)(24.305 g/mol) +
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g $\text{O}_2 \rightarrow$ 80.608 g MgO

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms Na + 2 molecules $\text{H}_2\text{O} \rightarrow$

2 formula units NaOH + 1 molecule H_2

2 mol Na + 2 mol $\text{H}_2\text{O} \rightarrow$

2 mol NaOH + 1 mol H_2

Mass:

2Na : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$: (4 mol)(1.008 g/mol) +
(2 mol)(15.999 g/mol) = 36.030 g

2NaOH : (2 mol)(22.990 g/mol) +
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)
= 79.994 g

H_2 : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g $\text{H}_2\text{O} \rightarrow$
79.994 g NaOH + 2.016 g H_2

82.01 g reactants = 82.01 g products

Supplemental Practice Problems Solutions Manual Chemistry

Julia Burdge



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Phillips, McGraw-Hill Education, 1996-06 Ebook: Chemistry Julia Burdge, 2014-10-16 Chemistry Third Edition by Julia

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Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry

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