

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a <u>proton donor</u>	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid	HCl	MINERAL ACIDS	They are derived from mineral salts i.e. chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3
Sulphuric acid	H_2SO_4		
Nitric acid	HNO_3		

Other mineral acids known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Derived from sulphites	Ethanoic acid (CH_3COOH)	CITRIC ACID from lemons
Carbonic acid	H_2CO_3	Derived from carbonates		TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Derived from phosphates	Methanoic acid ($HCOOH$)	ACETIC ACID from vinegar
Nitrous acid	HNO_2	Derived from nitrites		LACTIC ACID from sour milk
				Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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Study Guide Acid And Bases

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Table of Contents Study Guide Acid And Bases

1. Understanding the eBook Study Guide Acid And Bases
 - The Rise of Digital Reading Study Guide Acid And Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Acid And Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Acid And Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Acid And Bases
 - Personalized Recommendations
 - Study Guide Acid And Bases User Reviews and Ratings
 - Study Guide Acid And Bases and Bestseller Lists
5. Accessing Study Guide Acid And Bases Free and Paid eBooks

- Study Guide Acid And Bases Public Domain eBooks
 - Study Guide Acid And Bases eBook Subscription Services
 - Study Guide Acid And Bases Budget-Friendly Options
6. Navigating Study Guide Acid And Bases eBook Formats
- ePub, PDF, MOBI, and More
 - Study Guide Acid And Bases Compatibility with Devices
 - Study Guide Acid And Bases Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Study Guide Acid And Bases
 - Highlighting and Note-Taking Study Guide Acid And Bases
 - Interactive Elements Study Guide Acid And Bases
8. Staying Engaged with Study Guide Acid And Bases
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Acid And Bases
9. Balancing eBooks and Physical Books Study Guide Acid And Bases
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Acid And Bases
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Acid And Bases
- Setting Reading Goals Study Guide Acid And Bases
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Acid And Bases
- Fact-Checking eBook Content of Study Guide Acid And Bases
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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