

### Sample Problem 1:

Calculate the *Molarity* of a solution made by dissolving 2.355 g of copper (II) acetate (FW 181.5) in 50.0 mL of water.

$$0.259504 \text{ M} = 0.260 \text{ M}$$

KEY

$$\frac{2.355 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{50.0 \text{ mL soln}} \times \frac{1 \text{ mol Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{181.5 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2} \times \frac{1000 \text{ mL soln}}{1 \text{ L soln}} = 0.259504 \text{ M} = 0.260 \text{ M}$$

38F

### Sample Problem 2:

Calculate the *concentration in ppm, pp thousand* of a solution made by dissolving 2.355 g of copper (II) acetate (FW 181.5) and diluting up with water in a 50.0 mL volumetric flask. The density of solution at the given temperature is 1.0013 g/mL.

$$4.70388 \times 10^4 \text{ ppm} ; 4.70 \times 10^4 \text{ ppth}$$

$$\frac{2.355 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{50.0 \text{ mL soln}} \times \frac{1 \text{ mL soln}}{1.0013 \text{ g soln}} = \frac{x}{1.0 \times 10^6}$$
$$\frac{0.0470388 \text{ g Cu}(\text{C}_2\text{H}_3\text{O}_2)_2}{1 \text{ g soln}} = \frac{x}{1.0 \times 10^6} ; \quad x = 47,000 \text{ ppm}$$

OR

47.0388 ppthousand

### Sample Problem 3:

How many grams of solute would be used to prepare 1.50 L of a 0.250 M NaOH solution?

$$1.50 \text{ L NaOH soln} \times \frac{0.250 \text{ mol NaOH}}{1 \text{ L NaOH soln}} \times \frac{40.0 \text{ g NaOH}}{1 \text{ mol NaOH}} = 15.0 \text{ g NaOH}$$

dissolved in approximately  
0.75 L, then  
brought up to total Volume of 1.52

$$\begin{array}{r} \text{Na} \quad 23 \\ \text{O} \quad 16 \\ \text{H} \quad 1 \\ \hline 40.0 \text{ g} \\ \text{mole} \end{array}$$

# Solution Concentration Practice Problems

**Lorraine Zentz**



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## **Table of Contents Solution Concentration Practice Problems**

1. Understanding the eBook Solution Concentration Practice Problems
  - The Rise of Digital Reading Solution Concentration Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Solution Concentration Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solution Concentration Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Concentration Practice Problems
  - Personalized Recommendations
  - Solution Concentration Practice Problems User Reviews and Ratings
  - Solution Concentration Practice Problems and Bestseller Lists
5. Accessing Solution Concentration Practice Problems Free and Paid eBooks

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- Solution Concentration Practice Problems eBook Subscription Services
- Solution Concentration Practice Problems Budget-Friendly Options
- 6. Navigating Solution Concentration Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Solution Concentration Practice Problems Compatibility with Devices
  - Solution Concentration Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Solution Concentration Practice Problems
  - Highlighting and Note-Taking Solution Concentration Practice Problems
  - Interactive Elements Solution Concentration Practice Problems
- 8. Staying Engaged with Solution Concentration Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Solution Concentration Practice Problems
- 9. Balancing eBooks and Physical Books Solution Concentration Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Solution Concentration Practice Problems
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Concentration Practice Problems
  - Setting Reading Goals Solution Concentration Practice Problems
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution Concentration Practice Problems
  - Fact-Checking eBook Content of Solution Concentration Practice Problems
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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