

Solubility Rules Worksheet

1. Classify each substance as being soluble or insoluble in water.

- a. Potassium bromide – _____ f. Silver iodide - _____
b. Lead (II) carbonate – _____ g. Cadmium (II) sulfide - _____
c. Barium sulfate – _____ h. Zinc carbonate - _____
d. Zinc hydroxide – _____ i. Silver acetate - _____
e. Sodium acetate – _____ j. Copper (II) sulfide – _____

2. Identify the two new compounds which form if the solutions were mixed. Circle any compounds that would precipitate from the solution (are insoluble).

	KBr	Na_2CO_3	CaS	NH_4OH
AgNO_3				
BaCl_2				
$\text{Al}(\text{NO}_3)_3$				
CuSO_4				

3. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formu	Name	Solubility
$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$		
	Barium hydroxide	
FeCO_3		
NaOH		
	Rubidium nitrate	
MgSO_4		
	Zinc chloride	
	Zinc hydroxide	

Solubility Rules Lab Answer Key

Douglas Llewellyn

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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