

# SNC1D3 – Solubility Curve worksheet SOLUTIONS

For questions 1-5 the amount of solute is given, and a temperature is stated. If all the solute could be dissolved in 100 grams of water at the stated temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 1) 60g KCl at 70°C      supersaturated [above line]
- 2) 90g KNO<sub>3</sub> at 60°C      unsaturated [below line]
- 3) 110g NaNO<sub>3</sub> at 45°C      saturated [on line]
- 4) 10g KClO<sub>3</sub> at 10°C      supersaturated [above line]
- 5) 60g NH<sub>4</sub>Cl at 70°C      unsaturated [below line]

For questions 6-10 a solute and a temperature are given. Tell how many grams of each solute must be added to 100 grams of water to form a **saturated solution** (on the line) at that given temperature

- 6) NaNO<sub>3</sub> at 30°C      96g
- 7) KClO<sub>3</sub> at 70°C      36g
- 8) KNO<sub>3</sub> at 45°C      75g
- 9) KCl at 40°C      59g
- 10) NaCl at 90°C      40g

For questions 11-13, tell which solution is more concentrated (more solute dissolved)

- 11) At 50°C - A) a saturated solution of KNO<sub>3</sub> [85g]  
B) a saturated solution of NH<sub>4</sub>Cl [52g]
- 12) At 50°C - A) a saturated solution of KNO<sub>3</sub> [85g]  
B) an unsaturated solution of NaNO<sub>3</sub> consisting of 100g of the solute dissolved in 100g of water.
- 13) At 50°C - A) a saturated solution of NaNO<sub>3</sub> [116g]  
B) a supersaturated solution of NH<sub>4</sub>Cl consisting of 60g of the solute dissolved in 100g of water.

- 14) Which compound is most soluble at 40°C? KCl
- 15) Which compound is least soluble at 90°C? SO<sub>2</sub>
- 16) How many grams of KNO<sub>3</sub> will dissolve in 100g of water at 60°C? 108g
- 17) If 100g of a saturated solution of KNO<sub>3</sub> at 70°C is cooled to 20°C, how much KNO<sub>3</sub> will crystallize (in grams) 100g

- 18) If 200g of a saturated solution of KClO<sub>3</sub> at 50°C were evaporated to dryness, what would be left in the beaker?  
42g KClO<sub>3</sub>

OR:  $\frac{21}{135} = \frac{x}{200} \Rightarrow 100x = 21 \times 200 \Rightarrow 100x = 4200 \Rightarrow x = \frac{4200}{100} \Rightarrow x = 42$

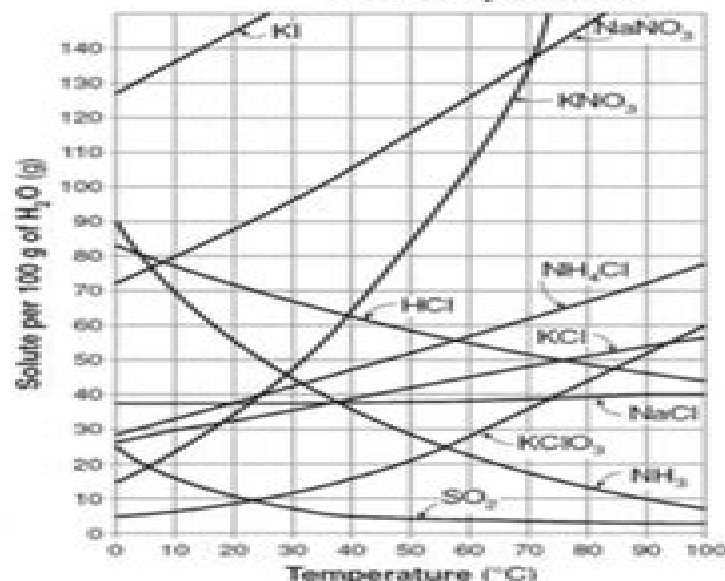
- 19) How much KCl will dissolve in 200g of water at 80°C 102g [51g x 2]
- 20) Which is more soluble at 30°C? KNO<sub>3</sub> or NaNO<sub>3</sub>? NaNO<sub>3</sub>
- 21) How much NH<sub>4</sub>Cl will dissolve in 50g of water at 60°C? 29g

- 22) A saturated solution of NH<sub>4</sub>Cl was made at 50°C in 100g of water. If the temperature is raised to 90°C, how much more NH<sub>4</sub>Cl must be added to be saturated? 20g [52g in 50°C and 72g in 90°C  $\Rightarrow 72 - 52 = 20$ ]

- 23) A saturated solution of KClO<sub>3</sub> at 70°C in 200g of water is cooled to 18°C. How much solute will precipitate out? 56g [36g at 70°C in 100g of water so 36x2=72 at 70°C in 200g of water  $\Rightarrow$  8g at 18°C in 100g of water, so 8x2=16 at 18°C in 200g of water] [72 - 16 = 56]

- 24) How much more SO<sub>2</sub> will dissolve in 50g of water at 10°C than at 40°C? 5.5g [16/2 at 10°C = 8g  $\Rightarrow$  5g/2 at 40°C = 2.5g] [8 - 2.5 = 5.5]

Solubility Curves



# Solubility Guidelines For Aqueous Solutions Worksheet Answers

**J Spring**



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