

PART ONE THERMODYNAMICS

CHAPTER 1: First Law of Thermodynamics

- ✓ 1.1 Assuming that the atmosphere is isothermal at 0 °C and that the average molar mass of air is 29 g mol⁻¹, calculate the atmospheric pressure at 20,000 ft above sea level.

SOLUTION

$$h = (2.0 \times 10^4 \text{ ft})(12 \text{ in ft}^{-1})(2.54 \text{ cm in}^{-1})(10^{-2} \text{ m cm}^{-1}) \\ = 6096 \text{ m}$$

$$P = P_0 e^{-gMh/RT}$$

$$P = (1.013 \text{ bar}) \exp \left[\frac{-(9.8 \text{ m s}^{-2})(29 \times 10^{-3} \text{ kg mol}^{-1})(6096 \text{ m})}{(8.314 \text{ J K}^{-1} \text{ mol}^{-1})(273 \text{ K})} \right] \\ = 0.472 \text{ bar}$$

- ✓ 1.2 Calculate the second virial coefficient of hydrogen at 0 °C from the fact that the molar volumes at 50.7, 101.3, 202.6, and 303.9 bar are 0.4634, 0.2386, 0.1271, and 0.09004 L mol⁻¹, respectively.

SOLUTION

$$\frac{PV}{RT} = 1 + \frac{B}{V} + \frac{C}{V^2} + \dots$$

P/bar	50.7	101.3	202.6	303.9
V/L mol ⁻¹	0.4634	0.2386	0.1271	0.09004
PV/RT	1.035	1.064	1.134	1.205
(1/V)/mol L ⁻¹	2.158	4.191	7.868	11.106

Solution Manual Chrupatla

Marco Cascella



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Table of Contents Solution Manual Chrupatla

1. Understanding the eBook Solution Manual Chrupatla
 - The Rise of Digital Reading Solution Manual Chrupatla
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Chrupatla
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Chrupatla
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Chrupatla
 - Personalized Recommendations
 - Solution Manual Chrupatla User Reviews and Ratings
 - Solution Manual Chrupatla and Bestseller Lists
5. Accessing Solution Manual Chrupatla Free and Paid eBooks
 - Solution Manual Chrupatla Public Domain eBooks
 - Solution Manual Chrupatla eBook Subscription Services
 - Solution Manual Chrupatla Budget-Friendly Options

6. Navigating Solution Manual Chrupatla eBook Formats
 - ePub, PDF, MOBI, and More
 - Solution Manual Chrupatla Compatibility with Devices
 - Solution Manual Chrupatla Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Manual Chrupatla
 - Highlighting and Note-Taking Solution Manual Chrupatla
 - Interactive Elements Solution Manual Chrupatla
8. Staying Engaged with Solution Manual Chrupatla
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Chrupatla
9. Balancing eBooks and Physical Books Solution Manual Chrupatla
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Chrupatla
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Manual Chrupatla
 - Setting Reading Goals Solution Manual Chrupatla
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solution Manual Chrupatla
 - Fact-Checking eBook Content of Solution Manual Chrupatla
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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