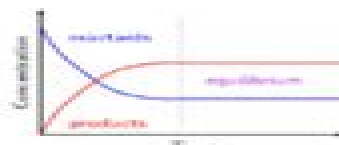


Equilibrium Practice Worksheet (Answers)

- The percent reaction of a reaction is 70%. Explain whether the equilibrium favours the reactants or products? What do you expect the magnitude of K to be? (Favours the products. Expect K value $\gg 0$.)
- Write an equilibrium constant expression for the reactions below.
 - $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ ($K = (\text{NH}_3)^2 / (\text{N}_2)(\text{H}_2)^3$)
 - $\text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightleftharpoons \text{BaSO}_4(\text{s}) + 2\text{NaNO}_3(\text{aq})$ ($K = (\text{NaNO}_3)^2 / (\text{Ba}(\text{NO}_3)_2)(\text{Na}_2\text{SO}_4)$)
- Determine the K_c for the reaction when 0.9 M of sulfur dioxide and 0.35 M of oxygen form 1.1 M of sulfur trioxide. ($K_c = 4.3$)
- Using the reaction from question 3. Determine which direction the equilibrium will shift if the Q value is 1×10^{-4} . (Will shift move toward the products)
- If the K_c value for an equilibrium is 3.8×10^{-6} , are the products or reactants favoured? What would the K_c value for the reverse reaction? (2.6×10^5)
- What would you expect the K value to be for the following graphs? (Would expect the first graph to have a K value of just above 1 and the second to have just below 1.)



- Determine which direction the equilibrium will shift when the following stresses are added.



- Temperature is increased - **right**
 - Pressure is decreased - **right**
 - Oxygen is added - **left**
 - A catalyst is added - **no shift**
 - Neon gas is added - **no shift**
- Graph the changes when for a, c and d. - (For a show increase in the products, decrease in reactants. For c should show an upward increase for O_2 and then a rate increase for reactants and rate decrease for products. For c should not see a change.)
 - Which direction would you expect the equilibrium to shift if NaOH was added to the reaction below. (Would shift right)

Skill Practice 45 Equilibrium Practice Answers

Rosina Ehmann



Skill Practice 45 Equilibrium Practice Answers:

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