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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar hcn h c n cf4 c f 4 cocl2
c o c l 2 nf3 n f 3 q3

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web the relative electronegativities of the bonded atoms is $\text{H} < \text{C} < \text{Cl}$ and so the bond moments all point toward the Cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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